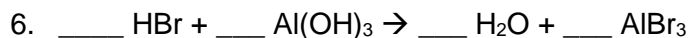
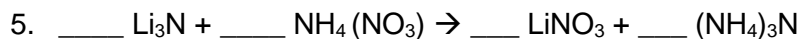
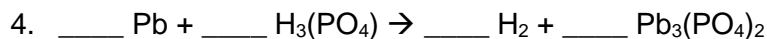
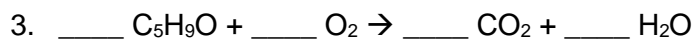
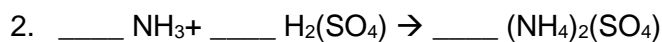
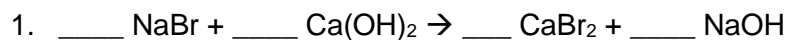
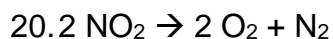
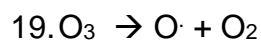
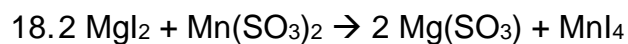
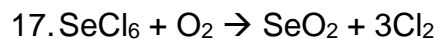
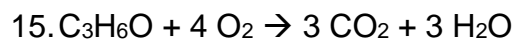
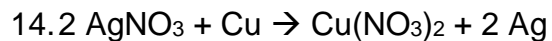
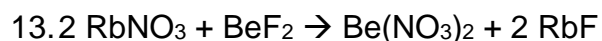
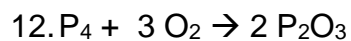
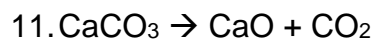
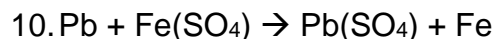
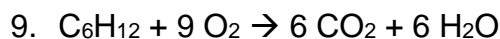
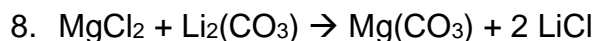
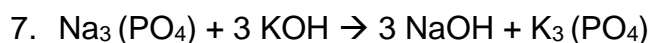


## Types of Chemical Reactions

Balance each of the following reactions and identify each type of reaction:



Identify each type of reaction:



## Types of Chemical Reactions Answers

Balance each of the following reactions and identify each type of reaction:

1.  $2 \text{NaBr} + \text{Ca}(\text{OH})_2 \rightarrow \text{CaBr}_2 + 2 \text{NaOH}$  **double displacement**
2.  $2 \text{NH}_3 + \text{H}_2\text{SO}_4 \rightarrow (\text{NH}_4)_2\text{SO}_4$  **synthesis**
3.  $4 \text{C}_5\text{H}_9\text{O} + 29 \text{O}_2 \rightarrow 20 \text{CO}_2 + 18 \text{H}_2\text{O}$  **combustion**
4.  $3 \text{Pb} + 2 \text{H}_3\text{PO}_4 \rightarrow 3 \text{H}_2 + \text{Pb}_3(\text{PO}_4)_2$  **single displacement**
5.  $\text{Li}_3\text{N} + 3 \text{NH}_4\text{NO}_3 \rightarrow 3 \text{LiNO}_3 + (\text{NH}_4)_3\text{N}$  **double displacement**
6.  $3 \text{HBr} + \text{Al}(\text{OH})_3 \rightarrow 3 \text{H}_2\text{O} + \text{AlBr}_3$  **double displacement**

Identify each type of reaction:

7.  $\text{Na}_3\text{PO}_4 + 3 \text{KOH} \rightarrow 3 \text{NaOH} + \text{K}_3\text{PO}_4$  **double displacement**
8.  $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2 \text{LiCl}$  **double displacement**
9.  $\text{C}_6\text{H}_{12} + 9 \text{O}_2 \rightarrow 6 \text{CO}_2 + 6 \text{H}_2\text{O}$  **combustion**
10.  $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$  **single displacement**
11.  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$  **decomposition**
12.  $\text{P}_4 + 3 \text{O}_2 \rightarrow 2 \text{P}_2\text{O}_3$  **synthesis**
13.  $2 \text{RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be}(\text{NO}_3)_2 + 2 \text{RbF}$  **double displacement**
14.  $2 \text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu}(\text{NO}_3)_2 + 2 \text{Ag}$  **single displacement**
15.  $\text{C}_3\text{H}_6\text{O} + 4 \text{O}_2 \rightarrow 3 \text{CO}_2 + 3 \text{H}_2\text{O}$  **combustion**
16.  $2 \text{C}_5\text{H}_5 + \text{Fe} \rightarrow \text{Fe}(\text{C}_5\text{H}_5)_2$  **synthesis**
17.  $\text{SeCl}_6 + \text{O}_2 \rightarrow \text{SeO}_2 + 3 \text{Cl}_2$  **single displacement**
18.  $2 \text{MgI}_2 + \text{Mn}(\text{SO}_3)_2 \rightarrow 2 \text{MgSO}_3 + \text{MnI}_4$  **double displacement**
19.  $\text{O}_3 \rightarrow \text{O} + \text{O}_2$  **decomposition**
20.  $2 \text{NO}_2 \rightarrow 2 \text{O}_2 + \text{N}_2$  **decomposition**