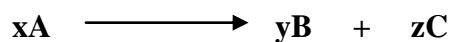
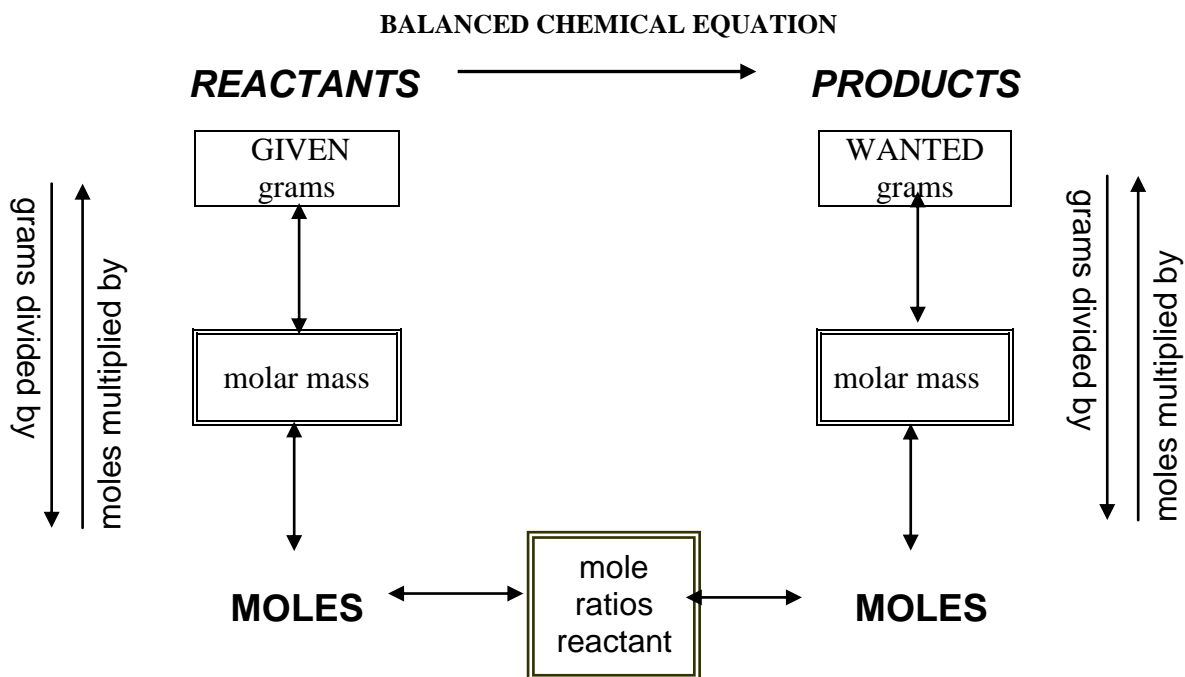


STOICHIOMETRY MAP FOR CHEMICAL REACTIONS

Double lined boxes are Conversion Factors to convert from one quantity to another.



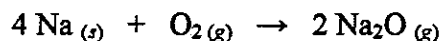
GIVEN:

WANTED:

$$\text{Grams A} \times \underbrace{\frac{1 \text{ mole A}}{\text{g A}}}_{\text{molar mass A}} \times \underbrace{\frac{y \text{ mole B}}{x \text{ mole A}}}_{\text{mole ratio from the balanced equation}} \times \underbrace{\frac{\text{g B}}{1 \text{ mole B}}}_{\text{molar mass B}} = \text{Gram B}$$

Stoichiometric Calculations

1. Sodium metal burns in air according to the balanced reaction shown below.



Complete the setups with the correct factors to answer the following questions:

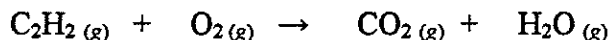
- (a) How many moles of oxygen are needed to completely react with 9.5 g of sodium?

$$\boxed{} \text{ g Na} \times \frac{1 \text{ mol Na}}{\boxed{} \text{ g Na}} \times \frac{\boxed{} \text{ mol O}_2}{\boxed{} \text{ mol Na}} = \boxed{} \text{ mol O}_2$$

- (b) How many grams of sodium are needed to produce 12.5 g of sodium oxide?

$$12.5 \text{ g Na}_2\text{O} \times \frac{1 \text{ mol Na}_2\text{O}}{62.0 \text{ g Na}_2\text{O}} \times \frac{\text{mol Na} \times \text{g Na}}{\text{mol Na}_2\text{O} \text{ mol Na}} =$$

2. Acetylene gas C_2H_2 undergoes combustion to form carbon dioxide and water when it is used in the oxyacetylene torch for welding. Balance the reaction and answer the following questions.



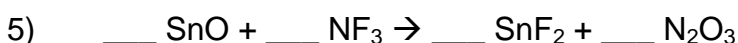
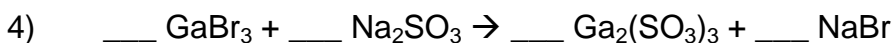
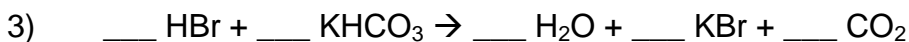
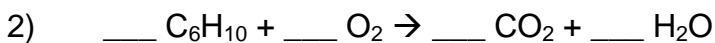
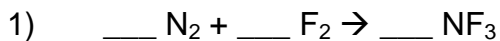
- (a) How many grams of water can form if 113 g of acetylene is burned?

- (b) How many grams of acetylene react if 1.10 mol of CO_2 are produced?

Stoichiometry Practice Worksheet

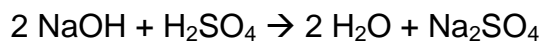
Balancing Equations and Simple Stoichiometry

Balance the following equations:



Solve the following stoichiometry grams-grams problems:

6) Using the following equation:



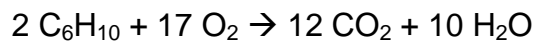
How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid?

7) Using the following equation:



How many grams of lithium nitrate will be needed to make 250 grams of lithium sulfate, assuming that you have an adequate amount of lead (IV) sulfate to do the reaction?

Use the following equation to answer questions 8-11:



- 8) If I do this reaction with 35 grams of C_6H_{10} and 45 grams of oxygen, how many grams of carbon dioxide will be formed?
- 9) What is the limiting reagent for problem 6? _____
- 10) How much of the excess reagent is left over after the reaction from problem 6 is finished?
- 11) If 35 grams of carbon dioxide are actually formed from the reaction in problem 6, what is the percent yield of this reaction?

